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↓ . D Change of Distance $S =$ Step 1:
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Balance the equations below: 1) $1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$ 2) $2 \text{ KClO}_3 \rightarrow 2 \text{ KCl} + 3 \text{ O}_2$ 3) $2 \text{ NaCl} + 1 \text{ F}_2 \rightarrow 2 \text{ NaF} + 1 \text{ Cl}_2$

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4) $2 \text{H}_2 + 1 \text{O}_2 \rightarrow 2 \text{H}_2\text{O}$ 5) $1 \text{Pb}(\text{OH})_2 + 2 \text{HCl} \rightarrow 2 \text{H}_2\text{O} + 1 \text{PbCl}_2$ 6) $2 \text{AlBr}_3 + 3 \text{K}_2\text{SO}_4 \rightarrow 6 \text{KBr} + 1 \text{Al}_2(\text{SO}_4)_3$ 7) $1 \text{CH}_4 + 2 \text{O}_2 \rightarrow 1 \text{CO}_2 + 2 \text{H}_2\text{O}$ 8) $1 \text{C}_3\text{H}_8 + 5 \text{O}_2 \rightarrow 3 \text{CO}_2 + 4 \text{H}_2\text{O}$ 9) $2 \text{C}_8\text{H}_{18}$ Answer key balancing chemical equations. . .

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